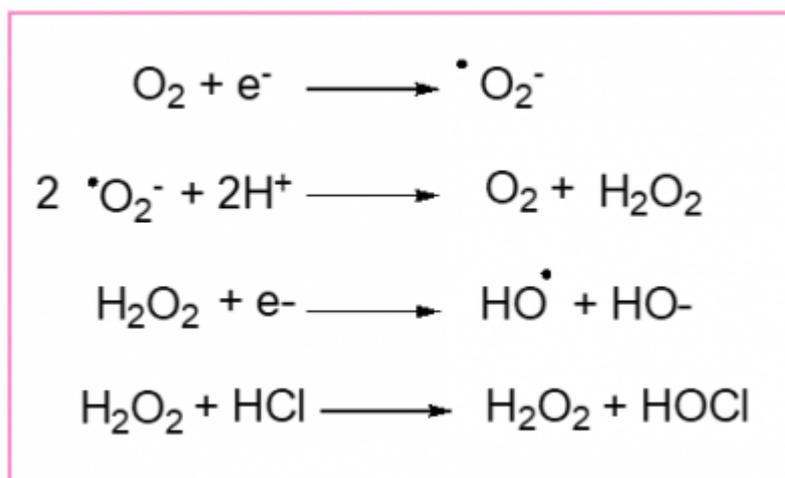


Reactive oxygen species

In its most stable state, oxygen exists in the form of dioxygen, O₂. What causes its possible toxicity is that this molecule can evolve into much more dangerous species: reactive oxygen species (ROS).

Initially O₂ can pick up an electron. A first reduced species is then formed, the superoxide anion ·O₂⁻. Because it comprises an unpaired electron, this anion is extremely reactive. To destroy it, today's living beings use specific enzymes, the superoxide dismutases (SOD), which will transform the superoxide anion into hydrogen peroxide, H₂O₂. But this hydrogen peroxide can also lead to a species containing an unpaired electron, the hydroxyl radical HO·.



Hydroxyl radicals react with many biological molecules. For example, they break down DNA strands and thus genetic information. They can also attack fatty acids to give new radicals, alkyl radicals, which will cause significant damage to cell membranes.

It's these species ·O₂⁻, H₂O₂, HO· and a few others like the hypochlorous acid HOCl, much more than O₂ itself (except in its triplet state which we have already mentioned), which made the period of the great oxidation so perilous for the bacteria and archaea of that time. They had to adapt and evolve (See [The dangerous liaisons of oxygen and life](#)). It is possible that it was this evolution that led to the formation of the first eukaryotic cells from modified archaea.

Articles from the Encyclopedia of the Environment describing the involvement of reactive oxygen species in the responses of living organisms to various environmental stresses:

- [The fixed life of plants](#)
- [How do plants cope with alpine stress?](#)
- [Cellular impact of solar UV](#)
- [What is the impact of air pollutants on vegetation?](#)
- [Human microbiotes: allies for our health](#)
- [Airborne particles: their health effects](#)
- [Spices and herbs: a benefit for our health?](#)

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