

# A carbon cycle disrupted by human activities

## Auteurs :

**JOYARD Jacques**, Directeur de recherche honoraire au CNRS, Laboratoire de Physiologie cellulaire et végétale, Université Grenoble Alpes

**SOMMERIA Joël**, Directeur de recherche CNRS, LEGI (Laboratoire des Ecoulements Géophysiques et Industriels), Université Grenoble-Alpes.

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*All living beings are built from carbon atoms. These are extracted from atmospheric CO<sub>2</sub> by plants, algae and certain bacteria, using solar energy: this is photosynthesis. The respiration and decomposition of living beings release this CO<sub>2</sub> back into the atmosphere. In addition to this short life cycle, there is a slow geological cycle that stores carbon in the form of limestone and fossil hydrocarbons. Limestone comes from the shells of marine organisms while hydrocarbons are formed by burial of organic sediments. The combustion of fossil resources currently represents a short circuit from this slow cycle to the short cycle that largely dominates natural regeneration processes. This leads to a rapid accumulation of CO<sub>2</sub> in the atmosphere, causing global warming, and ocean acidification that can disrupt marine life.*

## 1. The origins

**Carbon** is a fairly abundant element produced -like most light elements in the core of stars- by **nuclear fusion of hydrogen** (see [Nuclear Energy](#)). The carbon on Earth and the other planets comes from a star that exploded before the formation of the Solar System. During such an explosion (supernova) the condensable matter formed clouds of gas and dust which then aggregated to form planets in the early **Solar Nebula**. Like all non-radioactive elements [\[1\]](#), this initial carbon is permanently conserved and recycled since the formation of the Earth, passing through different environments and combining with different molecules.

This primitive material, prior to the formation of the Earth, has reached us almost intact in some meteorites called carbonaceous chondrites. It is also found in comets. Remarkably, carbon forms various elementary organic chemical compounds, such as **hydrocarbons**, alcohols and amino acids (see [How to study the organic molecules of comets](#)). Giant planets like Jupiter or

Saturn are composed mainly of hydrogen (90%) and helium (nearly 10%), but carbon is also present (0.1%) in the form of **methane** (CH<sub>4</sub>) or other hydrocarbons.

On rocky planets like the Earth, the high temperatures during their formation have led to the separation of the metallic elements (iron and nickel) that form the **core of** the less dense silicates that make up the **mantle**. It is estimated that carbon has been distributed in comparable proportions between nucleus [2] and mantle. In the latter, pure carbon subjected to high pressure and temperature crystallizes into diamonds: they are found on the surface in former highly active volcanic pipes, where rapid cooling has blocked any change in state. However, it is mainly present in the form of **carbonates** constituting between 0.003 and 0.03% of the mass of the mantle. It is released into the atmosphere as **carbon dioxide**, CO<sub>2</sub>, during volcanic eruptions. On Mars and Venus, volcanism also emits CO<sub>2</sub>, which is the main component of the atmosphere. CO<sub>2</sub> was also one of the main components of the Earth's atmosphere before **photosynthesis** appeared, which released oxygen while fixing atmospheric carbon as organic matter (see [The biosphere, a major geological player](#)).

## 2. Earth carbon reservoirs

Carbon is stored in different forms in media called **reservoirs**, the main ones are listed in Table 1. These reserves are expressed in gigatonnes of carbon (1GtC = 10<sup>12</sup> kg of carbon). More intuitively, these quantities can be reported on the total surface area of the globe (510 million km<sup>2</sup> or 510 x 10<sup>12</sup> m<sup>2</sup>), so that 1 kg/m<sup>2</sup> corresponds to a total of 510 GtC. The atmosphere currently contains 1.69 kg/m<sup>2</sup> of carbon in the form of CO<sub>2</sub> well mixed by the winds. This quantity can easily be deduced [3] from the atmospheric CO<sub>2</sub> concentration, currently 400 ppmv (one ppmv indicates a proportion of 10<sup>-6</sup> by volume, or the number of molecules). And a 1 ppmv increase corresponds to 2.12 GtC.

The ocean contains nearly 50 times more carbon than the atmospheric reservoir. The dissolved CO<sub>2</sub> reacts with water to give **carbonic acid** H<sub>2</sub>CO<sub>3</sub>, itself in equilibrium with the **hydrogen carbonate** ion HCO<sub>3</sub><sup>-</sup> (also called bicarbonate) and the **carbonate** ion CO<sub>3</sub><sup>2-</sup>. The HCO<sub>3</sub><sup>-</sup> form represents nearly 95% of ocean carbon. The ocean-atmosphere exchange transfers CO<sub>2</sub> either in one direction (atmosphere to ocean) or in the other (ocean to atmosphere), tending towards a temperature-dependent balance of concentration ratios.

The total mass of living oceanic organisms (called **ocean biomass**) contains a more limited amount of carbon, 6 g/m<sup>2</sup>, mainly in the form of plankton and bacteria. Terrestrial vegetation represents 1 kg/m<sup>2</sup> of carbon, while 4 kg/m<sup>2</sup> is stored in the organic debris of the soil (constituting **humus**). These are averages relative to the total surface area of the globe.

**Table 1. Carbon reservoirs** (mantle and terrestrial core are not mentioned due to high uncertainties and limited interactions with the environment).

Carbon repartition on Earth	Global stocks (Gigatons, Gt)	Global stock (1kg/m <sup>2</sup> = 510 Gt)
Atmosphere (in 2015)	860	1,69
Solubilized in oceans	38 000	76
Biomass from oceans	3	0,00
Land plants	500	1
Soil	2 000	4
Fossil carbon including: exploitable (coal, oil, natural gas)	10 000 000	20 000
	~2 000	4
Carbonate sedimentary rocks (limestone CaCO <sub>3</sub> & magnesite MgCO <sub>3</sub> )	70 000 000	140 000





*Figure 1. Limestone rocks near the summit of Mt Everest, formed by shells of marine organisms 500 million years ago [Source: Conrad Anker, <http://www.montana.edu/everest/resources/pdf/GearActivity/RocksofEverest.pdf>].*

However, a larger carbon reservoir is found in **calcareous** rocks, formed of **calcium carbonate** ( $\text{CaCO}_3$ ), combined with **magnesium** in variable proportions. The total mass of carbon immobilized in this way is estimated at about  $140 \text{ t/m}^2$ . As the mass of carbon represents 12% of the limestone, this represents about  $1000 \text{ t/m}^2$  of limestone, an average thickness of about 500 m. Limestone is a major geological formation such as the Paris Basin (especially in the form of chalk), the Colorado Grand Canyon, or large mountain ranges. They can be found as far as the summit of Mount Everest (Figure 1).

Another part of the carbon, estimated at  $20 \text{ t/m}^2$ , is stored as fossil carbon of organic origin (unoxidized). This carbon is found mainly in sediments and sedimentary rocks, in very low concentrations (kerogens). A small portion evolves into methane ( $\text{CH}_4$ ), which can be trapped on the ocean floor and in frozen soils as hydrates. Only a small fraction of the fossil carbon (less than 0.01%) has been concentrated, matured and sequestered to produce exploitable deposits of hydrocarbons (gas and oil) and coal (see [Oil: evidences for its biological origin](#)).

At time scales of hundreds of millions of years, sediments are carried away by tectonic plate movements and partially integrate into the mantle, whose total carbon reserve, very poorly known, is estimated to be even higher than that of limestone rocks.

### 3. Flows disrupted by human activity

The carbon flows between the different reservoirs are illustrated in Figure 2 (taken from [\[4\]](#)) which distinguishes between the pre-industrial situation and subsequent changes (red numbers). The first is the combustion of fossil fuels reserves, which produced  $7.8 \text{ GtC/year}$  in the decade 2000, taken as a reference for these data. The other numbers in red correspond to the system reactions that will be discussed in section 4.

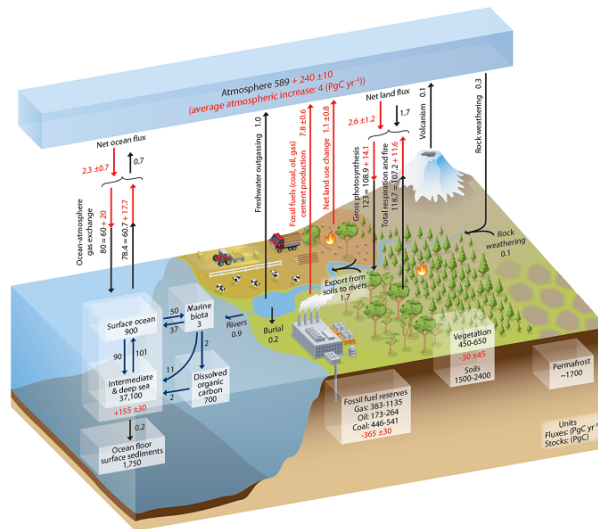


Figure 2. Representation of estimated carbon stocks on Earth that could support trade at the century scale (excluding limestone rocks), and annual carbon flows, in billions of tonnes of carbon (or petagrams, 1 PgC = 1 GtC). Black flows represent equilibrium "natural" trade, i. e. those that were the only ones that existed before the beginning of the industrial era, and black values for stocks represent the (reconstituted) state of 1750, before the beginning of industrial activities. Red flows represent the effect of human activities (increased, decreased or created), for the average of the decade 2000-2009, and red values for stocks represent the change from 1750 to 2011. [Source IPCC 2013, ref [4].]

CO<sub>2</sub> is permanently exchanged between the atmosphere and the surface ocean, by dissolution or degassing. Degassing is dominant in equatorial regions, where CO<sub>2</sub>-rich deep water rises and warms, and dominant dissolution in polar regions, due to surface water cooling. At mid-latitudes, seasonal temperature variations control the direction of exchanges. The two effects, dissolution and degassing, are approximately balanced on average with an overall flow of about 80 GtC/year in each direction.

The most important flow is related to photosynthesis{ind-text} Bioenergetic process that allows plants, algae and some bacteria to synthesize organic matter from CO<sub>2</sub> in the atmosphere using sunlight. Solar energy is used to oxidize water and reduce carbon dioxide in order to synthesize organic substances (carbohydrates). The oxidation of water leads to the formation of O<sub>2</sub> oxygen found in the atmosphere. Photosynthesis is the basis of autotrophy, it is the result of the integrated functioning of the chloroplast within the cell and organism. {end-tooltip} of terrestrial vegetation, capturing about 100 GtC/year, theoretically able to extract all CO<sub>2</sub> from the atmosphere in less than a decade. This is also known as **primary production**, which supplies organic matter to the entire earth's life. This flow is almost exactly compensated by the CO<sub>2</sub> emissions due to the breathing{ind-text} This refers both to the gaseous exchanges resulting from the inhalation and exhalation of air by living organisms (CO<sub>2</sub> carbon dioxide release and O<sub>2</sub> oxygen absorption) and to cellular respiration which, by degrading glucose with oxygen, allows energy to be obtained.{end-tooltip}, including that of the plants themselves and the bacteria that break down organic matter, as well as fires (e.g. forest fires); (read [Peatlands and marshes, remarkable wetlands](#)).

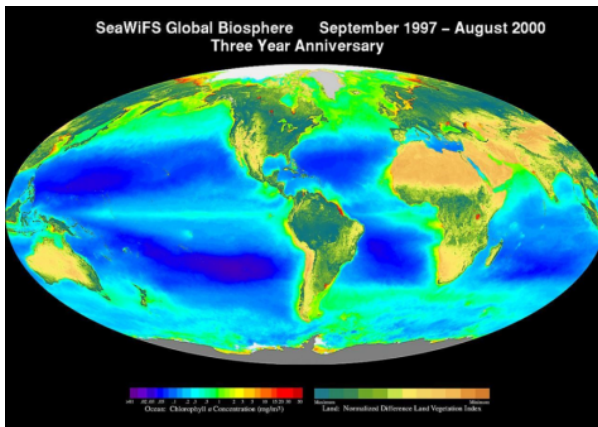


Figure 3. Distribution of the biosphere on Earth. Combination of views taken over 8 years by the SeaWiFS instrument on board the OrbView-2 spacecraft. The colours represent (a) marine surfaces: quantity of chlorophyll (in mg/m<sup>3</sup>, representing planktonic biomass, left-hand coloured scale), (b) land surfaces: vegetation index (right-hand coloured scale). [Source: NOAA].

The ocean is also an important medium for primary photosynthetic production, mainly by microorganisms that constitute what is

called **phytoplankton**: microalgae (including diatoms), cyanobacteria,... Satellites have provided accurate knowledge of population concentration, distribution and renewal rates and have provided credible estimates of primary production in the oceans. Observation of the chlorophyll content of the oceans (Figure 3) shows that areas between the 40<sup>th</sup> and 60<sup>th</sup> parallels in the North and South of the planet are richer than tropical areas (with the exception of the outlet of large rivers such as the nutrient-rich Amazon). It is possible to estimate primary oceanic production from such chlorophyll concentration observations, leading to an overall figure of 50 GtC per year, or about 50% of terrestrial production. It is remarkable that the carbon stored in marine life is nearly 200 times lower than that in terrestrial environments (3 GtC versus 500 GtC). Phytoplankton that dominate marine biomass store little carbon due to their short life span, unlike trees that dominate terrestrial biomass.

This primary production varies greatly in the ocean, regionally and seasonally, as phytoplankton are extremely sensitive to environmental fluctuations. Thus, the phosphate and nitrate richness of some waters (e.g. at the mouths of some rivers) locally fertilizes the surface layer of the ocean and leads to explosions in planktonic productivity (we speak of plankton blooms). This increase in photosynthesis reduces the CO<sub>2</sub> content of surface water, which is compensated by a dissolution of atmospheric CO<sub>2</sub>. Phytoplankton and algae are grazed by consumer organisms (zooplankton, fish, molluscs) that regenerate most of the CO<sub>2</sub> through their respiration and decomposition.

In addition, dead organisms or faecal particles carry organic carbon deep down where they feed living organisms (called heterotrophs) that remain dependent on surface photosynthesis. However, there is a notable exception near **volcanic fumaroles** where life is maintained by **chemosynthesis**, a reaction similar to photosynthesis where solar energy is replaced by the chemical energy of hydrogen sulphide (see [Microbes in extreme environments](#) and the focus on [Black smokers' ecosystems](#)).

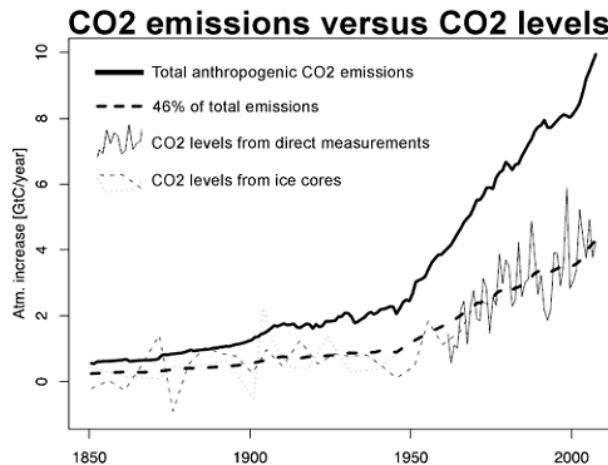
Only a small proportion of the marine organic carbon produced at the surface, about 0.2 GtC per year, accumulates in ocean sediments and slowly decomposes there without oxygen, forming hydrocarbons on a scale of millions of years. Coal is of terrestrial origin, its production began with the appearance of the first large trees during the Carboniferous era, 300 to 350 million years ago (see [The biosphere, a major geological player](#)).

Carbonated shells of marine organisms sediment and dissolve at great depth under pressure, releasing carbonate ions. This dissolution enriches deep waters with carbonates, such as sedimentation and recycling of organic particles ("biological pump"). In contrast, in shallow basins, these shells accumulate and slowly form limestone that is entrained in geological processes. Together with organic sediments, they are the only oceanic carbon sinks for the ocean-atmosphere system. In return, a CO<sub>2</sub> flux of about 0.1 GtC/year is re-emitted by **volcanoes** into the atmosphere.

The dissolution (**alteration**) of the limestone supplies the ocean with carbonates for 0.1 Gt/year (this dissolution, as well as that of silicates, also results in atmospheric CO<sub>2</sub> dissolved in rainwater for a total of 0.3 GtC/year). This carbon passes through rivers, joining that caused by soil erosion. The ocean receives 0.9 GtC/year from rivers, but this input is almost compensated by an estimated 0.7 GtC/year of ocean degassing in the pre-industrial situation.

These exchanges between atmosphere, rocks and living beings have played a considerable role in the ancient history of the Earth (see [The biosphere, a major geological player](#)), leading to significant variations in atmospheric composition and climate. However, these natural geological flows remain low: they are typically a thousand times lower than those of photosynthesis and nearly a **hundred times lower** than those of fossil fuel combustion. Over the past century, it is **human activities** that control the transformations of our environment.

## 4. The accumulation of CO<sub>2</sub> over the last century



*Figure 4. Annual production of CO<sub>2</sub> by burning fossil resources and changes in land use (deforestation) since the beginning of the industrial era, compared to the annual increase in measured atmospheric CO<sub>2</sub> (both expressed in Gt of carbon). This comparison shows that only half (46%) of the CO<sub>2</sub> emitted accumulates in the atmosphere. Concentration measurements prior to 1960 are obtained from bubbles trapped in ice. [Source W. Knorr, 2009, ref. [5]]*

There has been a rapid increase in atmospheric CO<sub>2</sub> content since the beginning of the industrial era, following an increase from a level that remained almost constant at around 280 ppmv for at least the previous thousand years. Not only is the atmospheric content increasing, but its **rate of increase** is also accelerating, as shown in Figure 4. This rate is currently about 4 GtC (or 2 ppmv) per year. If it remained constant, such a rate would increase the atmospheric content of the current value from 400 ppmv to 560 ppmv over 80 years, twice the pre-industrial value. This rate of increase is well correlated with emissions from human activities, as shown in Figure 4, mainly from the combustion of fossil resources (coal, oil and gas), but also includes deforestation and the production of cement that releases carbon from limestone (3% of CO<sub>2</sub> emissions). However, the total increase in atmospheric CO<sub>2</sub> remains half of the accumulation of these **anthropogenic** emissions, indicating that the other half is being reabsorbed, as we will see below.

This accumulation of atmospheric CO<sub>2</sub> induces **global warming** by increasing the **greenhouse effect** (radiation and climate link). It is also accompanied by **ocean acidification** that affects marine organisms.

The major contribution of fossil fuel combustion to the increase in atmospheric CO<sub>2</sub> is confirmed by the **isotopic** composition of CO<sub>2</sub> carbon (see [Radioactivity and nuclear reactions](#)). Carbon consists mainly of the isotope <sup>12</sup>C (containing 6 neutrons in addition to the 6 protons), but also of a small proportion of <sup>13</sup>C (7 neutrons) and <sup>14</sup>C (8 neutrons). Photosynthesis captures <sup>13</sup>C less efficiently than <sup>12</sup>C so that biologically derived carbon has a lower isotopic ratio of <sup>13</sup>C/<sup>12</sup>C than mineral carbon emitted by volcanoes. However, we observe that the isotopic ratio <sup>13</sup>C/<sup>12</sup>C of atmospheric CO<sub>2</sub> decreases, which marks its **biological origin**.

<sup>14</sup>C is produced by the nuclear reaction of cosmic radiation on nitrogen in the upper atmosphere, then integrated into CO<sub>2</sub> and captured during photosynthesis, so that it is incorporated into all living beings in known proportions (10<sup>-12</sup>). As it is radioactive, its concentration decreases with a half-life of 5730 years (this is what is used for dating biological remains). It has therefore completely disappeared into much older fossil fuels.

However, there is a decrease in the atmospheric <sup>14</sup>C concentration (after adjusting for the contributions of the nuclear explosions of the 1960s, which doubled the <sup>14</sup>C concentration). This shows that the increase in CO<sub>2</sub> comes from **fossil biological sources** [\[6\]](#) rather than from current biological sources.

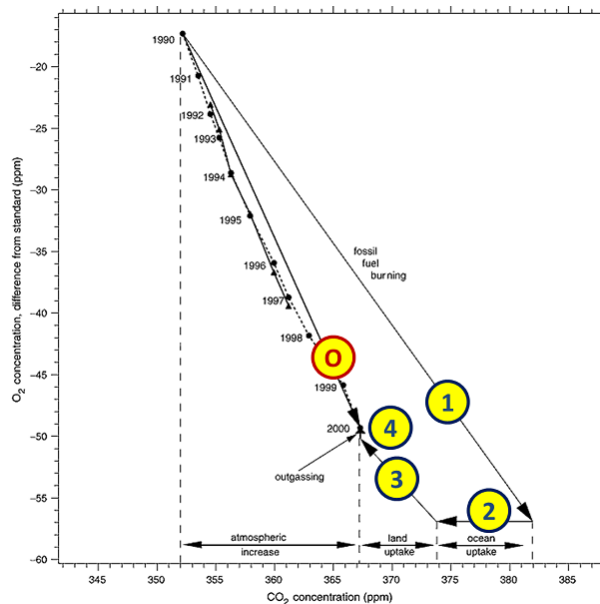


Figure 5. Decrease in atmospheric O<sub>2</sub> oxygen content as a function of the increase in CO<sub>2</sub> content during the 1990s (arrow 0). The arrow (1) represents the contribution of fossil fuel combustion, combining CO<sub>2</sub> increase and O<sub>2</sub> decrease in known proportions. The arrow (2) represents the dissolution in the ocean that reduces atmospheric CO<sub>2</sub> without variation of O<sub>2</sub>, while the arrow (3) represents photosynthesis, which decreases CO<sub>2</sub> by releasing O<sub>2</sub> in equal proportion. Finally, the ocean releases a small part of its dissolved oxygen, so there is no variation in CO<sub>2</sub> (arrow 4). [Source: IPCC Report 2001 (See ref [7])].

In addition, the combustion of fossil resources consumes oxygen. The resulting decrease in atmospheric O<sub>2</sub> concentration remains small but quite measurable. This decrease is shown in Figure 5 as a function of the increase in CO<sub>2</sub> concentration measured during the 1990s (15 ppmv in ten years). The measured increase in CO<sub>2</sub> is about half of that expected from the combustion of fossil resources (arrow 1). O<sub>2</sub> consumption is also lower, but not in the same proportion. These results make it possible to identify and evaluate **two causes of absorption of the CO<sub>2</sub> emitted**: **dissolution in the ocean** (arrow 2), which does not change the concentration of O<sub>2</sub>, and **photosynthesis**, which releases O<sub>2</sub> in the same proportion as it reduces CO<sub>2</sub> (arrow 3). A final, very limited effect corresponding to the degassing of dissolved oxygen due to ocean warming (arrow 4). In summary, of the 30 ppmv produced by fossil fuels in the 1990s, Figure 5 shows that 8 ppmv dissolved in the ocean, 7 ppmv were recycled by photosynthesis, while the remaining 15 ppmv accumulated in the atmosphere.

Over the decade 2000 shown in Figure 2, these figures translated into GtC (1 ppmv for 2.12 GtC, see section 2), increased to 7.8 GtC per year produced by fossil fuels, of which 2.3 GtC dissolved in the ocean and 1.5 GtC recycled by photosynthesis. The remainder, 4 GtC per year, has accumulated in the atmosphere.

In addition to CO<sub>2</sub>, carbon is also emitted into the atmosphere in the form of **methane** by anaerobic fermentation (away from the air) and by leaks from oil or gas fields. Although 200 times less concentrated than CO<sub>2</sub>, its contribution to the increase in the greenhouse effect is 25% of that of CO<sub>2</sub> [4]. Its impact on the greenhouse effect is about 25 times more powerful than CO<sub>2</sub> and its concentration in the atmosphere has increased by a factor of 2.5 compared to the pre-industrial situation. A sudden release of methane from frozen soils or marine sediments could accelerate global warming. However, the methane emitted into the atmosphere only persists for about ten years because it is slowly oxidized to CO<sub>2</sub>. Carbon is finally emitted in many other forms by human activities, in particular plastics produced at a rate of 0.3 Gt per year, of which an estimated one-tenth reaches the ocean and accumulates in the heart of vast eddies called gyres (see [Plastic pollution at sea: the seventh continent](#)).

## 5. Dissolution in the ocean

CO<sub>2</sub> dissolves in the surface ocean, tending to reach a concentration of equilibrium proportional to its partial pressure in the atmosphere. This is a general property of gas solubility called **Henry's Law**. This is why the dissolved CO<sub>2</sub> escapes when you open a bottle of champagne, following the drop in pressure. This degassing is increased by an increase in temperature, which reduces solubility. CO<sub>2</sub> thus escapes into the atmosphere in the warm tropical ocean, and on the contrary dissolves in the polar regions, the overall balance being neutral in a stationary regime. Global warming, on the other hand, induces global degassing. In the current period, however, it is the dissolution induced by the increase in atmospheric CO<sub>2</sub> that dominates.

The surface ocean currently contains mineral carbon at 24 ppm by mass [9] (24 g per tonne of water). If we take the total mass of the ocean,  $1.4 \times 10^{18}$  tonnes, this would represent a reserve of  $33 \times 10^{18}$  g of carbon, or 33,000 GtC. The reserve is estimated at 38,000 GtC because the concentration is a little higher at depth. At equilibrium, the concentration of dissolved CO<sub>2</sub> increases in proportion to its atmospheric pressure, in accordance with Henry's law, but 95% of the dissolved carbon is in the form of carbonate and hydrogen carbonate ions. Thus the relative increase in dissolved carbon is only about 1/10 [10] of the relative



increase in atmospheric CO<sub>2</sub>: an annual increase in this CO<sub>2</sub> of 2 ppmv over 400 ppmv, or 0.5%, therefore implies a 10 times smaller increase in dissolved carbon, or 0.05% of 24 ppm, giving 0.012 ppm. This carbon is only absorbed by the surface ocean: if it were absorbed by the entire ocean, the flux would be 17 GtC per year, whereas it is estimated at 2.3 GtC/year (Figure 2). The absorbed CO<sub>2</sub> is only transferred very slowly at depth.

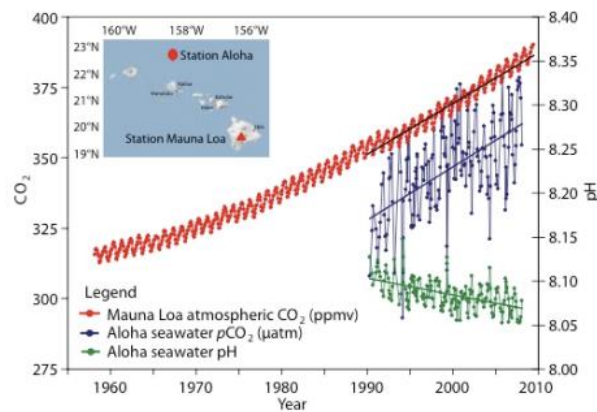


Figure 6. CO<sub>2</sub> content of the atmosphere (expressed in ppmv equivalent to 10<sup>-6</sup> atmosphere of partial pressure), and in surface water in the Hawaii Islands (expressed in equilibrium partial pressure with the same unit). The acidity of the water (pH) is also represented (green curve, right scale). The annual oscillations of atmospheric CO<sub>2</sub> are due to seasonal variations in photosynthesis and terrestrial respiration. Fluctuations in the ocean are greater because the mixing is much less effective than in the atmosphere, but the trend is confirmed by averages at many ocean sites. [Source: Oceanography, see ref. [9]].

Rather than the dissolved carbon concentration, oceanographers often prefer to measure the **equilibrium pressure** of CO<sub>2</sub> in water. This is obtained by allowing a seawater sample to degas in an air volume small enough not to affect its concentration in the water. If this equilibrium pressure is higher than the partial pressure of CO<sub>2</sub> in the atmosphere, dissolved CO<sub>2</sub> will escape to restore the equilibrium, otherwise, on the contrary, atmospheric CO<sub>2</sub> will be dissolved. Figure 6 shows that this equilibrium pressure remains close to the partial pressure of atmospheric CO<sub>2</sub> and increases concomitantly, confirming the equilibrium of the surface ocean with atmospheric CO<sub>2</sub>.

In addition, the dissolution of CO<sub>2</sub> releases H<sup>+</sup> ions by reaction with water, which corresponds to an **increase in acidity**. This is measured by the pH, which is even smaller when the solution is acidic. This decrease is clearly visible in Figure 6. Acidification tends to limit the synthesis of the coral's calcareous skeleton, in a variable way depending on the species [11].

However, this dissolution equilibrium is limited to the ocean in direct contact with the atmosphere, a layer of a few hundred metres thick stirred by the wind. The deep ocean remains unaffected by **stratification**: warmer and therefore less dense water floats on the surface. Surface water penetrates deep only sporadically and very locally into the polar regions where, after concentration of salt by evaporation, it can exceed the density of the bottom due to its cooling. It is estimated that 50 x 10<sup>6</sup> m<sup>3</sup>/s is the surface water flow that feeds the deep ocean, compensated of course by an equivalent global flow of deep water rising to the surface (the total mass of the ocean remaining unchanged). This represents 1.5 x 10<sup>15</sup> tonnes per year, and the renewal of the deep ocean, 1.4 x 10<sup>18</sup> tonnes, takes place in about 1000 years.

Thus, each year, 1.5 x 10<sup>15</sup> tonnes of water in equilibrium with the current CO<sub>2</sub> concentration (380 ppmv for the decade 2000) is replaced by deep water in equilibrium with the pre-industrial concentration of 280 ppmv, *i.e.* 30% lower. As we have seen, this represents a difference of 10 times less dissolved carbon, *i.e.* 3% or 0.7 ppm. Multiplying by the total water flow, we obtain a well of 1 GtC/year buried towards the deep ocean.

In addition, we have seen that the concentration of dissolved carbon increases by 0.012 ppm per year in the surface ocean, in a sufficiently stirred layer to homogenize over the decade. By choosing as an indication a thickness of 300 m for this mixed layer (note: this layer is thicker than the "surface ocean" in Figure 2 which is limited to the area of light penetration allowing photosynthesis) over the entire ocean surface, 0.36 x 10<sup>15</sup> m<sup>2</sup>, the mass of water concerned is 1.1 x 10<sup>17</sup> tonnes, which leads to an annual stored carbon mass of 1.3 GtC/year in this mixed layer. Adding what is discharged to the deep ocean, 1 GtC/year, we find the 2.3 GtC/year deduced from atmospheric observations for the decade 2000. This global value of the absorbed flux is also confirmed by a mapping of the CO<sub>2</sub> fluxes measured between the atmosphere and the ocean [8].

This flux must increase in proportion to the atmospheric CO<sub>2</sub> concentration as long as warming does not affect ocean currents. However, this purely physical assessment is very schematic. It assumes constant flows of biological origin, linked to sedimentation and dissolution processes, which play a key role in the vertical mixing of dissolved carbon, with a flow of 11 GtC/year.



## 6. The influence of vegetation

Terrestrial vegetation plays a key role in CO<sub>2</sub> fluxes, as well as in the storage of carbon in the form of visible vegetation, but also in the form of humus in the soil. Thus grasslands, savannahs and even semi-desert regions contribute significantly to the overall stock, although their vegetation content is low, see Table 2 (see [Peatlands and marshes, remarkable wetlands](#)). It also shows that boreal forests store a large amount of carbon in the soil, while tropical forests have very abundant vegetation but soil that is not very capable of storing carbon.

**Table 2. Overall carbon reserves in vegetation and upper (1 m) soil** (Adapted from IPCC, 2001 [\[12\]](#)).

	Areas (10 <sup>6</sup> km <sup>2</sup> )	Global Carbon stocks (Gt C)		
Biome		Plants	Soil	Total
Tropical forests	17.6	212	216	428
Temperate forests	10.4	59	100	159
Boreal forests	13.7	88	471	559
Tropical savannas	22.5	66	264	330
Temperate grasslands	12.5	9	295	304
Deserts et semi-deserts	45.5	8	191	199
Tundra	9.5	6	121	127
Wetlands	3.5	15	225	240
Croplands	16.0	3	128	131
Total land area	151.2	466	2011	2477

CO<sub>2</sub> emissions from deforestation are now estimated at 1.1 GtC per year [\[13\]](#), contributing 25% to the rate of increase in greenhouse gases in the atmosphere. Deforestation has several origins: forest exploitation, the need for space for urban expansion, the conversion of forests into agricultural land, mining. For example, the boreal forest is destroyed over an area of 3,400 km<sup>2</sup> by open-pit mining of the Athabasca **oil sands** (Alberta, Canada) [\[14\]](#). **Biofuel** production by large scale agriculture is also devastating, promoting the development of monocultures, such as soya in the Amazon and palm oil in Indonesia, Malaysia and Thailand.

The rate of deforestation in tropical forests is of particular concern from both an ecological and a climatic point of view. It reaches 10,000 to 30,000 km<sup>2</sup> per year in the Amazon [\[15\]](#), 5% of the forest area per decade in South America, with similar values in Africa (see Table 3). In Southeast Asia, fires encouraged by the El Nino climate phenomenon are helping to reduce forest cover. In 1997, forest and peatland fires in Indonesia released between 0.8 and 2.5 GtC of CO<sub>2</sub> into the atmosphere [\[16\]](#). Stabilization of forest areas in Asia includes massive destruction of primary forests and their replacement by planted forests, resulting in a reduction of nearly a quarter of the carbon stored by vegetation in 15 years (Table 3): from 12-74 kg/m<sup>2</sup> in forests to 8.5-29 kg/m<sup>2</sup> in oil palm plantations [\[17\]](#).

**Table 3. Global deforestation in figures** (Adapted from IPCC, 2007 [\[18\]](#)).

Region	Forest area	Annual change (Thousands km <sup>2</sup> )		Carbo
		1990-2000	2000-2005	
	2005			1990
Africa	6,36	-44	-40	65
Asia	5,72	-8	+10	41
Europe and Russia	10,01	+9	+7	42
North and Central America	7,06	-3	-3	41
Oceania	2,06	-4	-4	12
South America	8,32	-38	-43	97
World	39,52	-89	-73	296

On the other hand, forests are spreading to other parts of the world, offsetting some of the deforestation. Inventories increased in Europe, North America and Oceania. In the mid-2000s, the overall net deforestation rate was about 70,000 km<sup>2</sup> per year, slightly lower than the rate of about 90,000 km<sup>2</sup> per year that prevailed [16] in the 1990s. This represents a relative loss of surface area of 2% per decade, and even more a loss of stored carbon of 3%, corresponding to 1.1 GtC per year (see Table 3).

However, the atmospheric balance of CO<sub>2</sub> and oxygen (see Figure 5) indicates that photosynthesis absorbs 1.5 GtC per year overall. The 1.1 GtC carbon source due to forest destruction is therefore more than offset by a 2.6 GtC carbon sink per year. This "missing carbon" is probably captured in soils where the overall measurement is even more difficult than for vegetation. Several factors can lead to an increase in photosynthesis. It is known that increasing atmospheric CO<sub>2</sub> concentration promotes photosynthesis if other environmental factors are favourable. The spread of nitrogen fertilizers and global warming could also encourage plant activity leading to increased storage of organic matter in the soil. The ocean could also contribute to this carbon sink, which would imply an increase in sedimentation from the estimate of 0.2 GtC per year shown in Figure 2, which corresponds to an average observed on older sediments.

## 7. The prospects for remediation of anthropogenic carbon emissions

As we have seen, the burning of fossil resources and deforestation lead to an accelerated accumulation of CO<sub>2</sub> in the atmosphere, a source of global warming and ocean acidification. Only 50% of these flows are reabsorbed by natural feedbacks. A reduction of at least 50% in emissions is thus necessary to hopefully stabilize the CO<sub>2</sub> concentration.

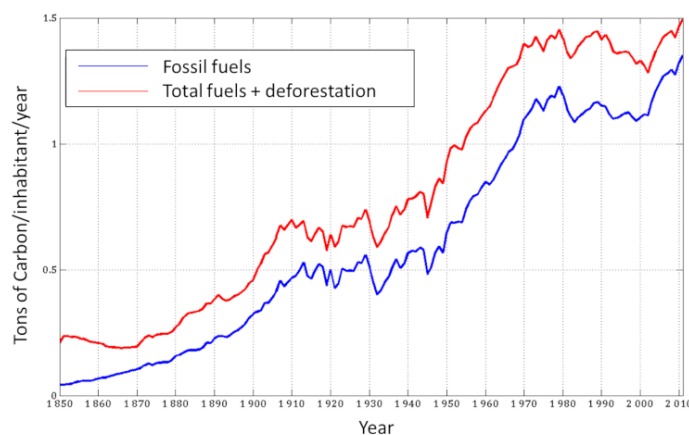


Figure 7. Evolution of annual CO<sub>2</sub> emissions per inhabitant (expressed in tonnes of carbon). Emissions from the combustion of fossil resources and cement production are given by the blue curve (data <http://cdiac.ornl.gov/>) while the red curve adds the effect of deforestation. [Source: Estimates provided by R. Houghton, corresponding to ref. [12].]

The increase in carbon emissions is primarily related to the increase in world population, which has increased fourfold in a

century (7.4 billion inhabitants in 2016 compared to about 1.8 billion in 1916). Per capita emissions have also increased overall by a factor of two over the same period, see Figure 7. The necessary reduction by a factor of two of emissions, in a context of high population growth, therefore requires a major change in technologies and consumption habits.

The per capita emission curve shows an inflection in the years 1970-2000 following the oil shock, the development of nuclear energy, and the decline of heavy industry (notably through the fall of the USSR). But there was a further increase in the 2000s following the massive use of coal, whose consumption increased by 50% in 10 years. It currently accounts for 30% of primary energy consumption (42% of CO<sub>2</sub> emissions from fossil fuels) and oil 36% (28% of emissions). The contribution of decarbonated energies remains low (nuclear 6% of primary consumption, hydroelectricity 6%, other renewable 1%) [18]. Nuclear energy is reluctant among the population, but has the advantage of a minimal influence on the natural environment, unlike most renewable energies. These remain dominated by conventional hydropower. Wind and solar energy, which are still marginal, are developing rapidly. But they require the development of storage techniques, or additional fossil fuels, to compensate for the intermittency of their activity. In addition, their impact is limited to electricity production, which accounts for hardly a third of the primary energy consumed.

Methods for deep CO<sub>2</sub> capture have been proposed, either in geological formations or on the ocean floor, by adapting mining and petroleum techniques. However, their long-term reliability is difficult to demonstrate and their energy consumption is considerable. Their development would thus only aggravate the problem of energy resources, leading to the exploitation of "unconventional" deposits that are increasingly devastating for the environment.

It is probably by limiting deforestation and making the best use of agricultural resources that the most effective gains can be achieved in the medium term. Agriculture occupies 1/3 of the ice-free land area (12% for crops, 22% for pastures), and is estimated to absorb 1/4 land primary production [19] (total organic matter production). Livestock is a source of 15% of greenhouse gas emissions, more than transport. Biofuels currently on the market (known as first generation) compete with food crops and cause deforestation. However, research on second- and third-generation biofuels raises hopes of using agricultural waste, improving current methanisation techniques (methane production by air-free fermentation).

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## References and notes

[1] However, it is necessary to mention the exception of Carbon 14 (constituting a proportion  $10^{-12}$  of the carbon of living beings). This carbon isotope is produced in the upper atmosphere by the impact of cosmic radiation on nitrogen, then is transformed back into nitrogen by radioactivity with a period of 5730 years.

[2] Chen *et al* (2014) *Hidden carbon in Earth's inner core revealed by shear softening in dense Fe<sub>7</sub>C<sub>3</sub>*. PNAS 111, 17755-17758, <http://www.pnas.org/content/111/50/17755.full.pdf>

[3] To do this, we notice that the total atmosphere contains 10.3 t/m<sup>2</sup>: it is the mass of the air column whose weight produces the average atmospheric pressure  $1.013 \times 10^5$  N/m<sup>2</sup>. The mole (N<sub>A</sub> molecules where N<sub>A</sub> is the Avogadro number) of CO<sub>2</sub> containing 12 g of carbon and 32 g of oxygen, or 44 g, while an atmospheric mole contains 29 g (weighted average between nitrogen and oxygen), the mass proportion of CO<sub>2</sub> is  $44/29 = 1.52$  times its proportion in molecules  $400 \times 10^{-6}$ , or  $607 \times 10^{-6}$  of 10.3 t/m<sup>2</sup>, equal to 6.25 kg/m<sup>2</sup> CO<sub>2</sub>. The mass proportion of carbon in CO<sub>2</sub> being  $12/44 = 27\%$ , we find that 400 ppmv represents 1.69 kg/m<sup>2</sup> of carbon or 860 GtC.

[4] IPCC Report 2013, chapter 6: Carbon and Other Biogeochemical Cycles, [http://www.ipcc.ch/report/ar5/wg1http://www.climatechange2013.org/images/report/WG1AR5\\_ALL\\_FINAL.pdf](http://www.ipcc.ch/report/ar5/wg1http://www.climatechange2013.org/images/report/WG1AR5_ALL_FINAL.pdf)

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[http://www.ideo.columbia.edu/res/pi/CO2/carbondioxide/pages/air\\_sea\\_flux\\_2010.html](http://www.ideo.columbia.edu/res/pi/CO2/carbondioxide/pages/air_sea_flux_2010.html).

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[10] This increase factor, also dependent on temperature, is called the "Reveal Factor", see Zeebe R. <http://www.eoearth.org/view/article/154468>

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[20] [https://en.wikipedia.org/wiki/Primary\\_production](https://en.wikipedia.org/wiki/Primary_production)

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